

Date Planned : __ / __ / __	Daily Tutorial Sheet - 15	Expected Duration : 90 Min
Actual Date of Attempt : __ / __ / __	Level - 3	Exact Duration : _____

- 157.** If Aufbau rule is not followed, K-19 will be placed in _____ block ?
(A) s **(B)** p **(C)** d **(D)** f
- *158.** Select equations having endothermic step:
(A) $S^{-}(g) \longrightarrow S^{2-}(g)$ **(B)** $Na^{+}(g) + Cl^{-}(g) \longrightarrow NaCl(s)$
(C) $N(g) \longrightarrow N^{-}(g)$ **(D)** $Al^{2+}(g) \longrightarrow Al^{3+}(g)$
- *159.** The correct order of radii is/are:
(A) $Pb > Pb^{2+} > Pb^{4+}$ **(B)** $In^{+} > Sn^{2+} > Sb^{3+} > Te^{4+}$
(C) $Co > Ni > Cu > Zn$ **(D)** $K^{+} > Li^{+} > Mg^{2+} > Al^{3+}$
- 160.** The electron affinity of chlorine is 3.7 eV. The energy released when 2.0 g of chloride is completely converted to Cl^{-} ion in gaseous phase is ($1\text{ eV} = 23.06\text{ kcal mol}^{-1}$).
(A) 4.8 kcal **(B)** 5.8 kcal **(C)** 6.8 kcal **(D)** 7.8 kcal
- 161.** The ground state electron configurations listed here are incorrect. Identify the mistakes that have been made in each and write the correct electronic configurations.
(A) Al : $1s^2 2s^2 2p^4 3s^2 3p^3$ **(B)** B : $1s^2 2s^2 2p^5$
(C) Cl : $1s^2 2s^2 2p^6 3s^2 3p^6$ **(D)** Cu : $1s^2 2s^2 2p^6 3s^2 3p^6 3d^9 4s^2$
- *162.** Which of the following elements have almost equal value of electronegativity?
(A) H **(B)** S **(C)** Te **(D)** P